

## Bis(2-chlorobenzoato- $\kappa$ O)bis(1-vinylimidazole- $\kappa$ N<sup>3</sup>)copper(II)

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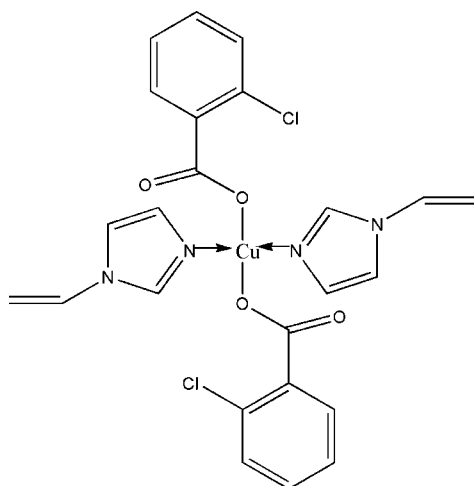
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.011$  Å;  $R$  factor = 0.073;  $wR$  factor = 0.192; data-to-parameter ratio = 13.7.

In the title compound,  $[\text{Cu}(\text{C}_7\text{H}_4\text{ClO}_2)_2(\text{C}_5\text{H}_6\text{N}_2)_2]$ , each  $\text{Cu}^{\text{II}}$  ion, located on an inversion center, has a slightly distorted square-planar coordination geometry formed by two 1-vinylimidazole molecules [ $\text{Cu}-\text{N} = 1.954$  (6) Å] and two 2-chlorobenzoate anions [ $\text{Cu}-\text{O} = 1.958$  (6) Å]. Weak intermolecular  $\text{C}-\text{H}\cdots\text{O}$  hydrogen bonds contribute to the crystal packing stability.

### Related literature

A square-planar coordination environment of  $\text{Cu}^{\text{II}}$  was also observed in bis(3-hydroxybenzoato- $\kappa$ O)bis(1*H*-imidazole- $\kappa$ N<sup>3</sup>)copper(II), see: Liu *et al.* (2006).



### Experimental

#### Crystal data

$[\text{Cu}(\text{C}_7\text{H}_4\text{ClO}_2)_2(\text{C}_5\text{H}_6\text{N}_2)_2]$   
 $M_r = 562.89$   
Monoclinic,  $P2_1/c$   
 $a = 7.9360$  (16) Å  
 $b = 11.236$  (2) Å  
 $c = 14.190$  (3) Å  
 $\beta = 104.36$  (3)°

$V = 1225.8$  (5) Å<sup>3</sup>  
 $Z = 2$   
Mo  $K\alpha$  radiation  
 $\mu = 1.15$  mm<sup>-1</sup>  
 $T = 293$  (2) K  
 $0.20 \times 0.10 \times 0.10$  mm

#### Data collection

Bruker SMART 1K CCD area-detector diffractometer  
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2004)  
 $T_{\text{min}} = 0.803$ ,  $T_{\text{max}} = 0.894$

2204 measured reflections  
2115 independent reflections  
1620 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.039$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.073$   
 $wR(F^2) = 0.192$   
 $S = 1.04$   
2115 reflections  
154 parameters

49 restraints  
H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.73$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.89$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C1}-\text{H1A}\cdots\text{O2}^{\text{i}}$	0.93	2.56	3.484 (10)	174
$\text{C3}-\text{H3A}\cdots\text{O1}^{\text{ii}}$	0.93	2.49	2.918 (8)	108
$\text{C5}-\text{H5A}\cdots\text{O2}^{\text{i}}$	0.93	2.45	3.342 (9)	160
$\text{C11}-\text{H11A}\cdots\text{O2}^{\text{iii}}$	0.93	2.60	3.460 (9)	155

Symmetry codes: (i)  $x - 1, y, z$ ; (ii)  $-x, -y, -z$ ; (iii)  $-x, y - \frac{1}{2}, -z + \frac{1}{2}$ .

Data collection: *SMART* (Bruker, 2001); cell refinement: *S SAINT* (Bruker, 2001); data reduction: *S SAINT*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL* and local programs.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CV2449).

### References

- Bruker (2001). *SMART* and *S SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.  
Liu, J.-W., Zhu, B. & Ng, S. W. (2006). *Acta Cryst.* **E62**, m3514–m3515.  
Sheldrick, G. M. (2004). *SADABS*. University of Göttingen, Germany.  
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

**supplementary materials**

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## Bis(2-chlorobenzoato- $\kappa O$ )bis(1-vinylimidazole- $\kappa N^3$ )copper(II)

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### Comment

In the title compound, (I) (Fig. 1), each Cu ion is coordinated by a pair of 1-vinylimidazole ligands and a pair of monodentate carboxylate groups, affording a square planar  $N_2O_2$  coordination geometry. The  $CuN_2O_2$  core involving the central atoms is almost perfectly square planar. The *trans* angles are all  $180^\circ$  for symmetry requirements and the *cis* ones are  $89.52(19)^\circ$  and  $90.48(19)^\circ$  for  $N-Cu-O$ , respectively. The  $Cu-N(\text{imidazole})$  distance is  $1.954(6)\text{ \AA}$  and The  $Cu-O$  bond distance is  $1.958(4)\text{ \AA}$ . These bond distances are comparable with the reported data (Liu *et al.*, 2006). The five atoms of  $CuN_2O_2$  are coplanar. Distances and angles in 1-vinylimidazole are normal. The weak intermolecular  $C-H\cdots O$  interactions (Table 1) stabilize the structure.

### Experimental

Copper(II) acetate hydrate(2.00 g, 10 mmol), 1-vinylimidazole(0.99 g, 10 mmol) and 2-chlorobenzoic acid(1.55 g, 10 mmol) were dissolved in water(40 ml). The pH of the solution was adjusted to 7 with 0.2M sodium hydroxide. The solution was filtered; blue single crystals of (I) were isolated after several days.

### Refinement

H atoms were positioned geometrically ( $C-H = 0.93\text{ \AA}$ ) and allowed to ride on their parent atoms with  $U_{\text{iso}}(H) = 1.2 U_{\text{eq}}(C)$ .

### Figures

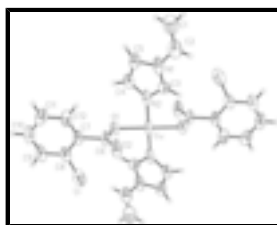


Fig. 1. The molecular structure of (I), showing 50% probability displacement ellipsoids and the atom-numbering scheme. The unlabelled atoms are related with the labelled ones by symmetry operation  $(-x, -y, -z)$ .

## Bis(2-chlorobenzoato- $\kappa O$ )bis(1-vinylimidazole- $\kappa N^3$ )copper(II)

### Crystal data

$[Cu(C_7H_4ClO_2)_2(C_5H_6N_2)_2]$

$M_r = 562.89$

Monoclinic,  $P2_1/c$

Hall symbol:  $-P\ 2ybc$

$F(000) = 574$

$D_x = 1.525\text{ Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\text{ \AA}$

Cell parameters from 25 reflections

# supplementary materials

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$a = 7.9360 (16) \text{ \AA}$	$\theta = 10\text{--}14^\circ$
$b = 11.236 (2) \text{ \AA}$	$\mu = 1.15 \text{ mm}^{-1}$
$c = 14.190 (3) \text{ \AA}$	$T = 293 \text{ K}$
$\beta = 104.36 (3)^\circ$	Block, blue
$V = 1225.8 (5) \text{ \AA}^3$	$0.20 \times 0.10 \times 0.10 \text{ mm}$
$Z = 2$	

## Data collection

Bruker SMART 1K CCD area-detector diffractometer	2115 independent reflections
Radiation source: fine-focus sealed tube graphite	1620 reflections with $I > 2\sigma(I)$
Thin-slice $\omega$ scans	$R_{\text{int}} = 0.039$
Absorption correction: multi-scan (SADABS; Sheldrick, 2004)	$\theta_{\text{max}} = 25.2^\circ, \theta_{\text{min}} = 2.3^\circ$
$T_{\text{min}} = 0.803, T_{\text{max}} = 0.894$	$h = -9 \rightarrow 9$
2204 measured reflections	$k = 0 \rightarrow 13$
	$l = 0 \rightarrow 16$

## Refinement

Refinement on $F^2$	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.073$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.192$	H-atom parameters constrained
$S = 1.04$	$w = 1/[\sigma^2(F_o^2) + (0.07P)^2 + 6P]$
2115 reflections	where $P = (F_o^2 + 2F_c^2)/3$
154 parameters	$(\Delta/\sigma)_{\text{max}} < 0.001$
49 restraints	$\Delta\rho_{\text{max}} = 0.73 \text{ e \AA}^{-3}$
	$\Delta\rho_{\text{min}} = -0.89 \text{ e \AA}^{-3}$

## Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

## Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
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Cu	0.0000	0.0000	0.0000	0.0439 (4)
Cl	0.3736 (3)	-0.36078 (18)	0.02395 (17)	0.0794 (6)
O1	0.0066 (5)	-0.1740 (3)	0.0077 (3)	0.0455 (10)
N1	-0.3796 (8)	0.0695 (5)	0.1425 (4)	0.0550 (14)
C1	-0.6165 (13)	0.1373 (8)	0.2115 (6)	0.088 (3)
H1A	-0.6754	0.0655	0.1961	0.105*
H1B	-0.6638	0.1972	0.2422	0.105*
O2	0.1925 (7)	-0.1406 (4)	0.1509 (3)	0.0647 (14)
N2	-0.1911 (7)	0.0008 (4)	0.0642 (4)	0.0527 (14)
C2	-0.4664 (11)	0.1540 (7)	0.1896 (5)	0.066 (2)
H2A	-0.4120	0.2270	0.2063	0.079*
C3	-0.2436 (8)	0.0933 (6)	0.1068 (5)	0.047
H3A	-0.1907	0.1676	0.1114	0.057*
C4	-0.3049 (9)	-0.0881 (6)	0.0741 (5)	0.0504 (15)
H4A	-0.3010	-0.1654	0.0513	0.060*
C5	-0.4234 (9)	-0.0489 (6)	0.1212 (5)	0.0537 (16)
H5A	-0.5138	-0.0917	0.1360	0.064*
C6	0.1057 (9)	-0.2068 (5)	0.0889 (5)	0.0501 (16)
C7	0.1031 (8)	-0.3403 (5)	0.1088 (4)	0.0433 (13)
C8	0.2177 (9)	-0.4153 (6)	0.0813 (4)	0.0513 (15)
C9	0.2146 (11)	-0.5370 (6)	0.0999 (6)	0.0662 (19)
H9A	0.2926	-0.5882	0.0812	0.079*
C10	0.0945 (11)	-0.5801 (6)	0.1461 (6)	0.0676 (19)
H10A	0.0906	-0.6613	0.1580	0.081*
C11	-0.0196 (11)	-0.5057 (7)	0.1749 (5)	0.0648 (18)
H11A	-0.0990	-0.5360	0.2071	0.078*
C12	-0.0162 (10)	-0.3853 (6)	0.1558 (5)	0.0580 (17)
H12A	-0.0943	-0.3344	0.1746	0.070*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cu	0.0696 (7)	0.0204 (5)	0.0350 (5)	0.0026 (5)	0.0001 (5)	0.0014 (4)
Cl	0.1031 (15)	0.0558 (11)	0.0838 (14)	0.0230 (11)	0.0319 (12)	0.0061 (10)
O1	0.060 (2)	0.030 (2)	0.043 (2)	0.0052 (19)	0.0066 (19)	0.0033 (18)
N1	0.088 (4)	0.033 (3)	0.035 (3)	0.009 (3)	-0.004 (3)	0.002 (2)
C1	0.115 (7)	0.071 (6)	0.079 (6)	0.004 (5)	0.029 (6)	-0.011 (5)
O2	0.094 (4)	0.029 (2)	0.058 (3)	-0.003 (2)	-0.006 (3)	-0.009 (2)
N2	0.075 (3)	0.024 (2)	0.048 (3)	0.008 (3)	-0.008 (3)	0.001 (2)
C2	0.096 (6)	0.054 (5)	0.042 (4)	0.004 (4)	0.004 (4)	0.007 (3)
C3	0.047	0.047	0.047	0.000	0.012	0.000
C4	0.068 (4)	0.035 (3)	0.043 (3)	0.008 (3)	0.006 (3)	0.003 (3)
C5	0.067 (4)	0.043 (3)	0.045 (4)	-0.001 (3)	0.001 (3)	0.009 (3)
C6	0.074 (4)	0.020 (3)	0.050 (4)	-0.004 (3)	0.004 (3)	0.000 (3)
C7	0.064 (3)	0.028 (3)	0.031 (3)	0.002 (2)	-0.001 (2)	-0.002 (2)
C8	0.072 (4)	0.037 (3)	0.040 (3)	0.011 (3)	0.004 (3)	-0.002 (3)
C9	0.091 (5)	0.039 (3)	0.062 (4)	0.017 (3)	0.008 (4)	-0.001 (3)
C10	0.091 (5)	0.035 (3)	0.063 (4)	-0.005 (3)	-0.007 (3)	0.007 (3)

## supplementary materials

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C11	0.086 (4)	0.053 (4)	0.051 (4)	-0.017 (3)	0.010 (3)	0.013 (3)
C12	0.085 (4)	0.042 (3)	0.047 (3)	-0.002 (3)	0.017 (3)	0.007 (3)

### *Geometric parameters (Å, °)*

Cu—N2 <sup>i</sup>	1.954 (6)	C3—H3A	0.9300
Cu—N2	1.954 (6)	C4—C5	1.356 (9)
Cu—O1 <sup>i</sup>	1.958 (4)	C4—H4A	0.9300
Cu—O1	1.958 (4)	C5—H5A	0.9300
Cl—C8	1.751 (7)	C6—C7	1.528 (8)
O1—C6	1.278 (7)	C7—C8	1.366 (8)
N1—C3	1.327 (8)	C7—C12	1.383 (9)
N1—C5	1.389 (9)	C8—C9	1.394 (10)
N1—C2	1.433 (9)	C9—C10	1.372 (11)
C1—C2	1.317 (10)	C9—H9A	0.9300
C1—H1A	0.9300	C10—C11	1.367 (11)
C1—H1B	0.9300	C10—H10A	0.9300
O2—C6	1.225 (7)	C11—C12	1.381 (9)
N2—C3	1.320 (8)	C11—H11A	0.9300
N2—C4	1.377 (8)	C12—H12A	0.9300
C2—H2A	0.9300		
N2 <sup>i</sup> —Cu—N2	180.0 (3)	C4—C5—N1	104.5 (6)
N2 <sup>i</sup> —Cu—O1 <sup>i</sup>	89.52 (19)	C4—C5—H5A	127.7
N2—Cu—O1 <sup>i</sup>	90.48 (19)	N1—C5—H5A	127.7
N2 <sup>i</sup> —Cu—O1	90.48 (19)	O2—C6—O1	125.6 (5)
N2—Cu—O1	89.52 (19)	O2—C6—C7	119.7 (6)
O1 <sup>i</sup> —Cu—O1	180.0 (4)	O1—C6—C7	114.6 (5)
C6—O1—Cu	109.9 (4)	C8—C7—C12	119.8 (6)
C3—N1—C5	107.1 (6)	C8—C7—C6	120.9 (6)
C3—N1—C2	125.1 (6)	C12—C7—C6	119.4 (6)
C5—N1—C2	127.8 (6)	C7—C8—C9	120.4 (7)
C2—C1—H1A	120.0	C7—C8—C1	120.9 (5)
C2—C1—H1B	120.0	C9—C8—C1	118.6 (5)
H1A—C1—H1B	120.0	C10—C9—C8	118.9 (7)
C3—N2—C4	103.6 (6)	C10—C9—H9A	120.5
C3—N2—Cu	125.9 (4)	C8—C9—H9A	120.5
C4—N2—Cu	130.4 (4)	C11—C10—C9	121.1 (7)
C1—C2—N1	125.6 (8)	C11—C10—H10A	119.4
C1—C2—H2A	117.2	C9—C10—H10A	119.4
N1—C2—H2A	117.2	C10—C11—C12	119.6 (7)
N2—C3—N1	113.2 (6)	C10—C11—H11A	120.2
N2—C3—H3A	123.4	C12—C11—H11A	120.2
N1—C3—H3A	123.4	C11—C12—C7	120.1 (7)
C5—C4—N2	111.6 (6)	C11—C12—H12A	119.9
C5—C4—H4A	124.2	C7—C12—H12A	119.9
N2—C4—H4A	124.2		
N2 <sup>i</sup> —Cu—O1—C6	92.5 (4)	Cu—O1—C6—O2	-4.0 (9)

N2—Cu—O1—C6	-87.5 (4)	Cu—O1—C6—C7	172.6 (4)
O1 <sup>i</sup> —Cu—N2—C3	-17.5 (5)	O2—C6—C7—C8	-92.3 (8)
O1—Cu—N2—C3	162.5 (5)	O1—C6—C7—C8	90.9 (7)
O1 <sup>i</sup> —Cu—N2—C4	160.1 (5)	O2—C6—C7—C12	87.2 (8)
O1—Cu—N2—C4	-19.9 (5)	O1—C6—C7—C12	-89.6 (7)
C3—N1—C2—C1	168.6 (8)	C12—C7—C8—C9	0.4 (10)
C5—N1—C2—C1	-8.3 (11)	C6—C7—C8—C9	179.9 (6)
C4—N2—C3—N1	0.2 (7)	C12—C7—C8—C1	-178.6 (5)
Cu—N2—C3—N1	178.3 (4)	C6—C7—C8—C1	0.9 (8)
C5—N1—C3—N2	-0.6 (7)	C7—C8—C9—C10	0.1 (10)
C2—N1—C3—N2	-178.0 (5)	C1—C8—C9—C10	179.0 (6)
C3—N2—C4—C5	0.3 (7)	C8—C9—C10—C11	-0.8 (11)
Cu—N2—C4—C5	-177.7 (4)	C9—C10—C11—C12	1.1 (11)
N2—C4—C5—N1	-0.7 (7)	C10—C11—C12—C7	-0.7 (11)
C3—N1—C5—C4	0.8 (7)	C8—C7—C12—C11	0.0 (10)
C2—N1—C5—C4	178.0 (6)	C6—C7—C12—C11	-179.6 (6)

Symmetry codes: (i)  $-x, -y, -z$ .

*Hydrogen-bond geometry* ( $\text{\AA}, ^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C1—H1A $\cdots$ O2 <sup>ii</sup>	0.93	2.56	3.484 (10)	174
C3—H3A $\cdots$ O1 <sup>iii</sup>	0.93	2.49	2.918 (8)	108
C5—H5A $\cdots$ O2 <sup>ii</sup>	0.93	2.45	3.342 (9)	160
C11—H11A $\cdots$ O2 <sup>iv</sup>	0.93	2.60	3.460 (9)	155

Symmetry codes: (ii)  $x-1, y, z$ ; (iii)  $-x, -y, -z$ ; (iv)  $-x, y-1/2, -z+1/2$ .

Fig. 1

